



2.13

Calculate the thermal conductivity of air, hydrogen, and carbon dioxide at 300 K, assuming ideal gas behavior. Compare your calculated values to values from Table A.4.

$$k = \frac{1}{\gamma - 1} \frac{c_v}{\pi d^2} \sqrt{\frac{m k_B T}{N \gamma}} \quad \text{eq. 2.12}$$

$$\gamma = \frac{c_p}{c_v}$$

$c_p$  heat capacity if pressure is const

$c_v$  heat capacity if volume is const

$$PV = nRT \quad R = 8.314 \frac{\text{J}}{\text{K mol}}$$

# Mayer's Relation

$$C_{pm} - C_{vm} = R$$

$M$  molar mass  $\frac{K_J}{K_{mol}}$

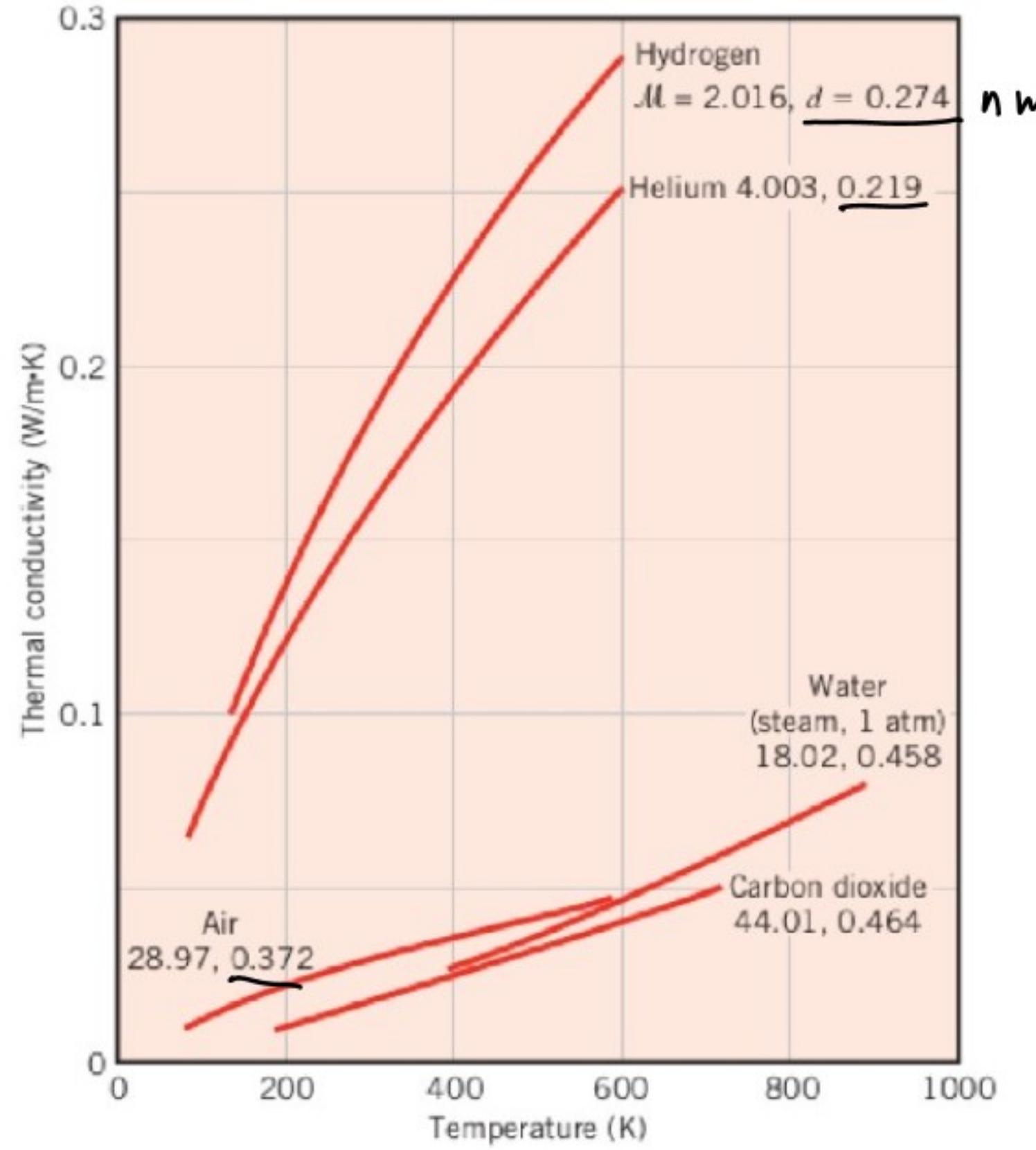
$$C_p \frac{KJ}{K, K}$$

$$C_{pm} \frac{KJ}{K_{mol} K}$$

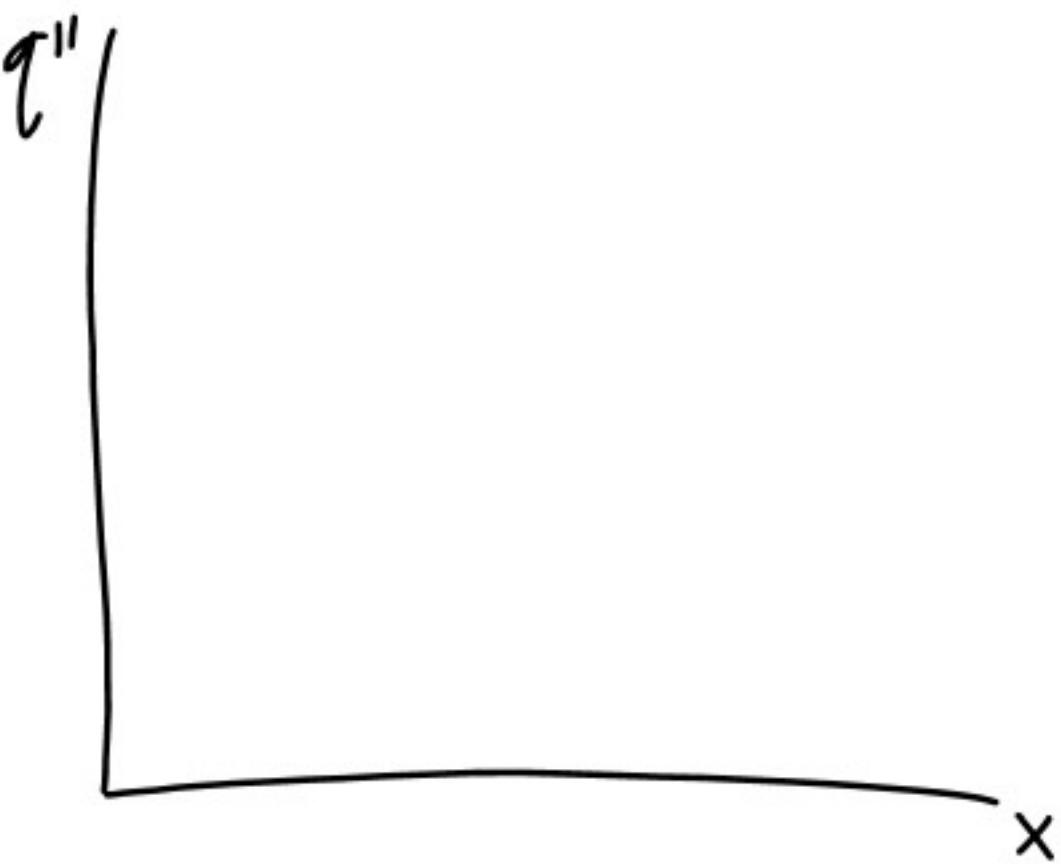
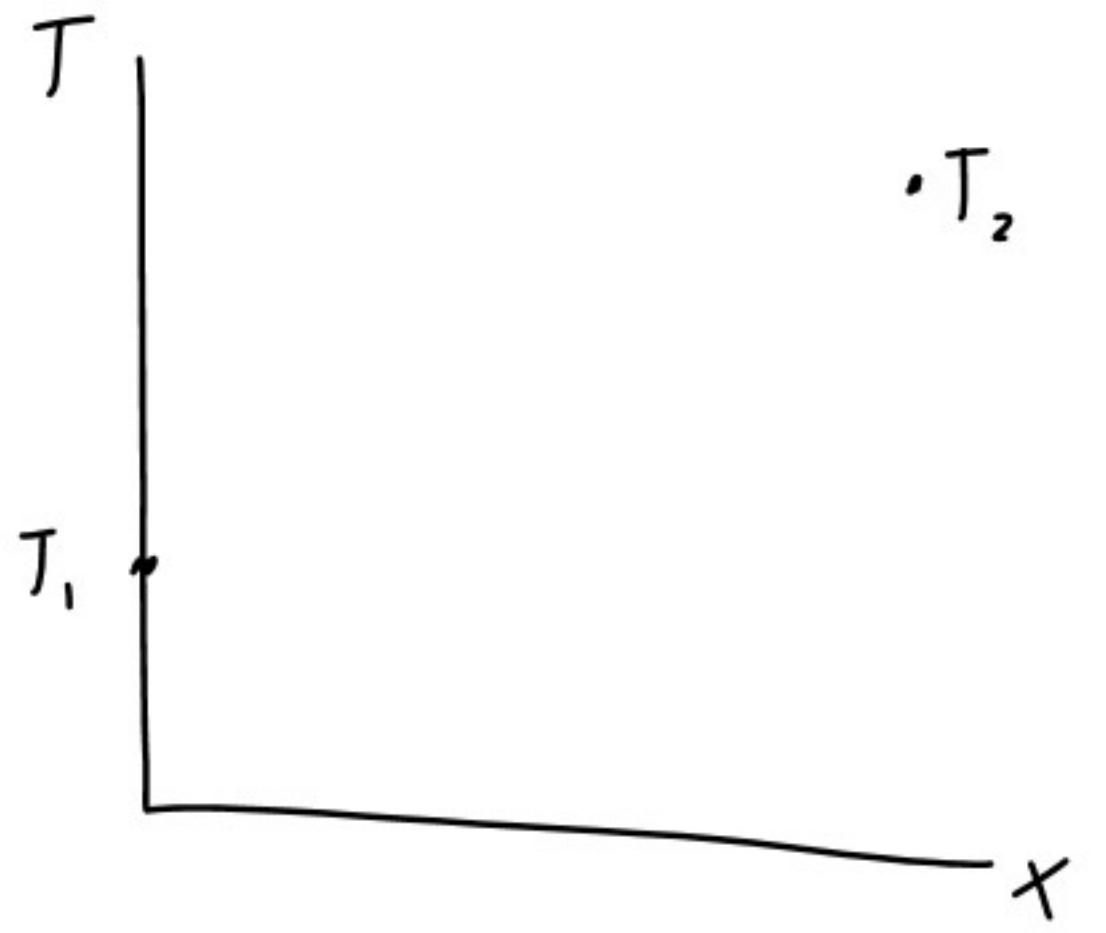
$$\frac{C_{pm}}{M} - \frac{C_{vm}}{M} = \frac{R}{M}$$

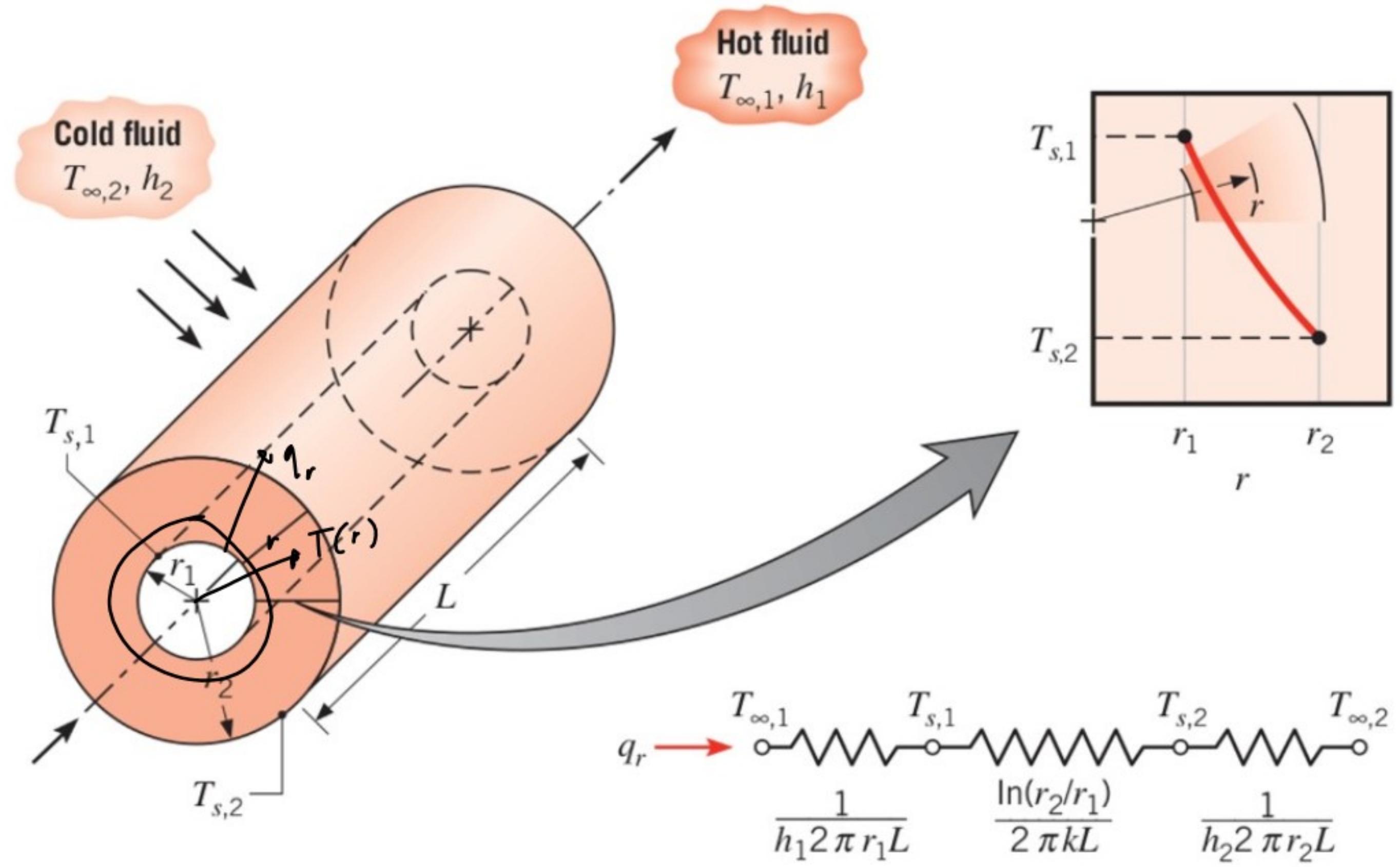
$$\boxed{C_p - C_v = \frac{R}{M}}$$

Fig 2.8



2.25





$$q_r = -kA \frac{dT}{dr} = -k(2\pi r L) \frac{dT}{dr}$$

$q''_r$  not constant

$$\frac{1}{r} \frac{\partial}{\partial r} \left( r \frac{\partial T}{\partial r} \right) + \frac{1}{r^2} \cancel{\frac{\partial^2 T}{\partial \theta^2}} + \cancel{\frac{\partial^2 T}{\partial z^2}} + \cancel{\frac{\partial T}{\partial t}} = \frac{1}{\alpha} \cancel{\frac{\partial T}{\partial t}}$$

0                    0                    0                    0  
no generation

$$\frac{1}{r} \frac{\partial}{\partial r} \left( r \frac{\partial T}{\partial r} \right) = 0$$

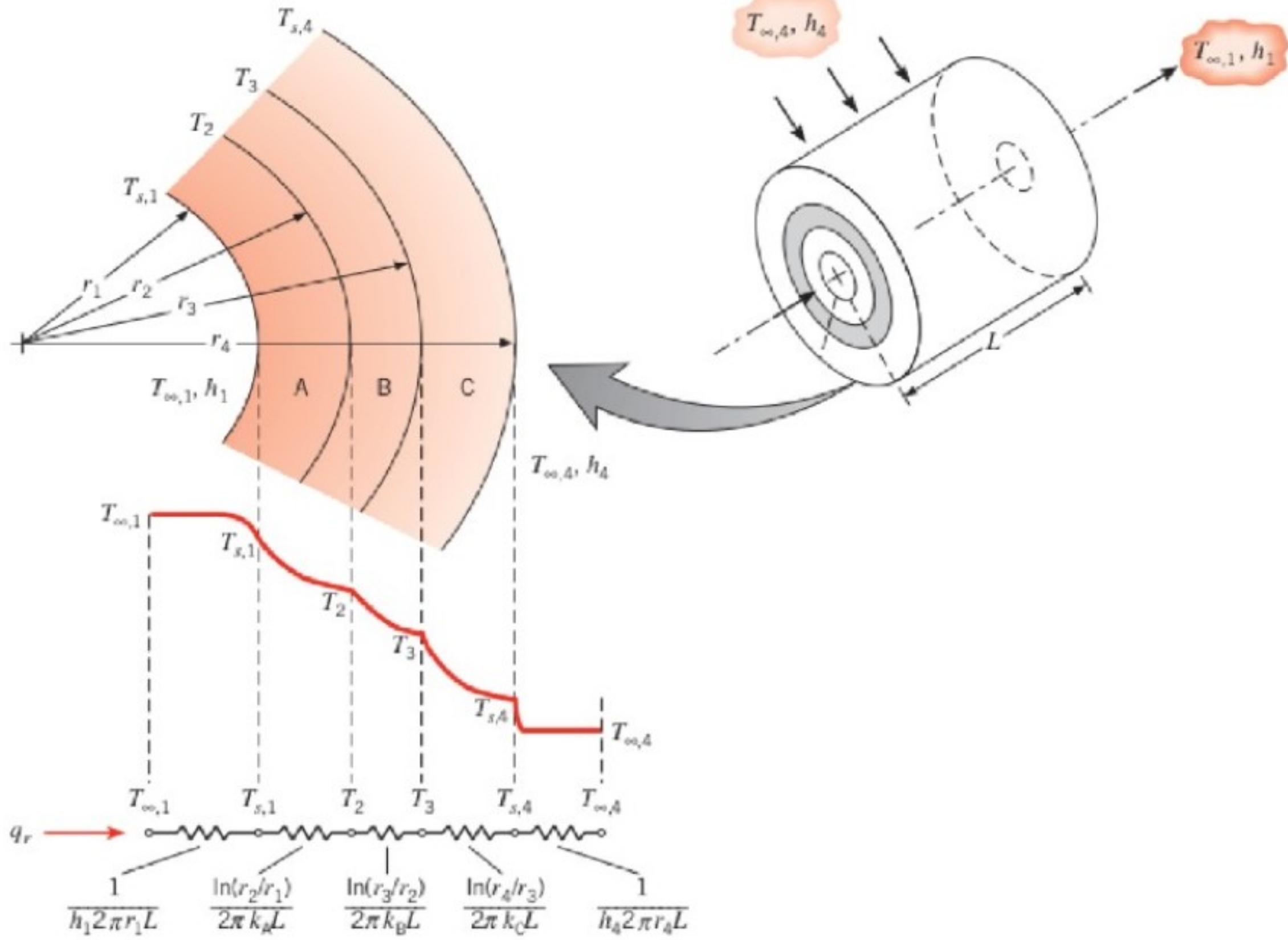
$$T(r) = C_1 \ln(r) + C_2$$

$$T(r_1) = T_{s1} \quad T(r_2) = T_{s2}$$

$$T(r) = \frac{T_{s1} - T_{s2}}{\ln(r_2/r_1)} \ln(r/r_1) + T_{s2}$$

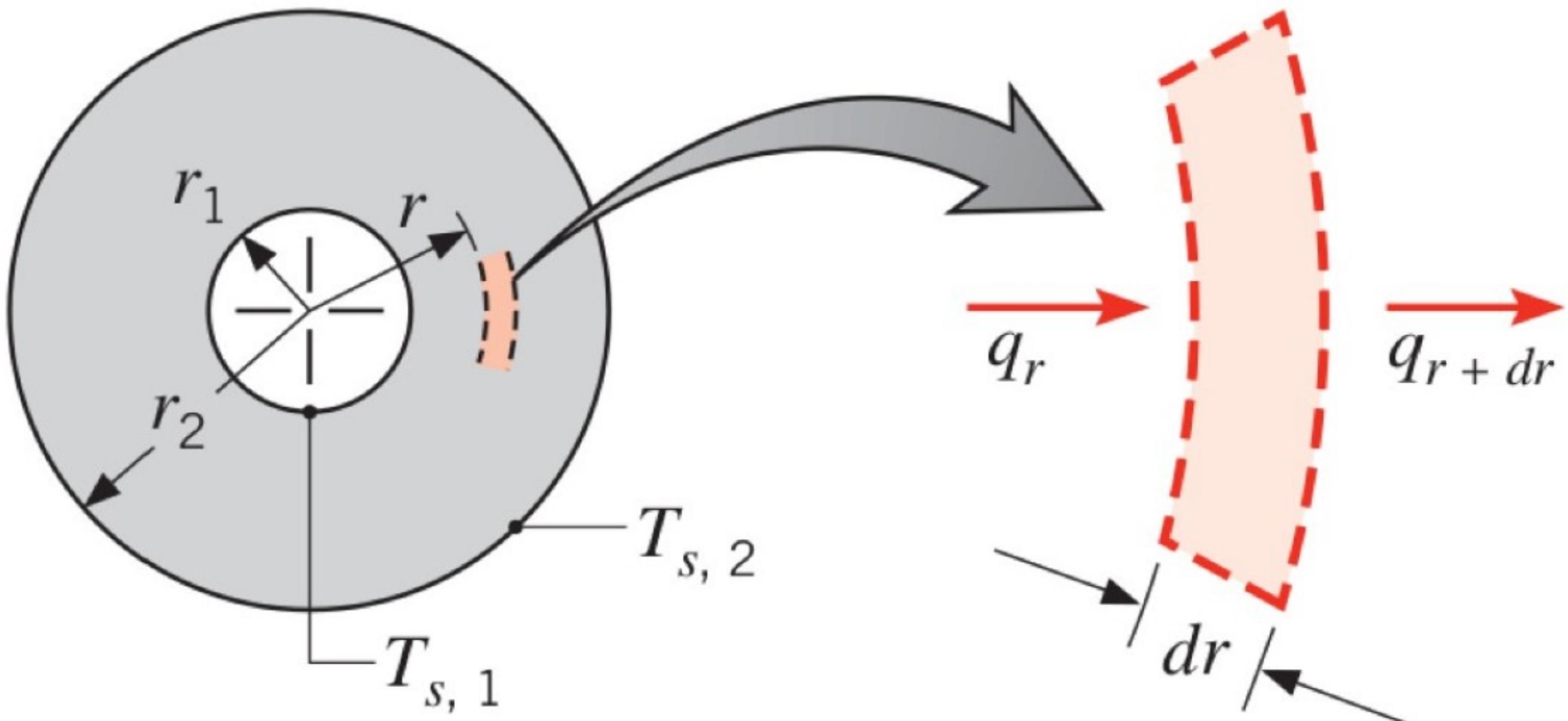
$$q_r = \frac{2\pi L k (T_{s1} - T_{s2})}{\ln(r_2/r_1)}$$

$$R_{cond} = \frac{\ln(r_2/r_1)}{2\pi L k}$$



1D

# Spherical Conduction



$$q_r = -k_A \frac{dT}{dr} = -k \cdot \pi r^2 \frac{dT}{dr}$$

$$\int \frac{q_r}{\pi r^2} dr = -k \int dT$$

$$\frac{q_r}{\pi r} \int_{r_1}^{r_2} \frac{1}{r^2} dr = -k \int_{T_{S1}}^{T_{S2}} dT$$

$$q_r = \frac{q \pi k (T_{S1} - T_{S2})}{(\frac{1}{r_1}) - (\frac{1}{r_2})}$$

$$R_{cond} = \frac{1}{q \pi k} \left( \frac{1}{r_1} - \frac{1}{r_2} \right)$$